Additions and Corrections

Contribution of Tunneling to Secondary Isotope Effects in Proton-Transfer Reactions [*J. Am. Chem. Soc.* **1984**, *106*, 2223–2224]. WILLIAM H. SAUNDERS, JR.

Page 2224, eq 5 should read: $1 - A^2 - 2B^2 - 3C^2 + A^2B^2 + 3B^2C^2 = D$.

Page 2224, paragraph 2: The effect of the above change is to make D vary within a given model for models 3–5. Model 3: $D = -0.21$ to -0.31. Model 4: $D = -0.32$ to -0.68. Model 5: $D = -0.37$ to -0.87. Models based on the correct equation given somewhat different numerical values, but none of the qualitative conclusions is changed.

Dimethylsilylene: Its Optical Absorption Spectrum and Reaction Kinetics [*J. Am. Chem. Soc.* **1984**, *106*, 7267–7268]. A. S. NAZRAN, J. A. HAWARI, D. GRILLER,* I. S. ALNAIMI, and W. P. WEBER*

The reported optical absorption spectrum assigned to dimethylsilylene ($\lambda_{max} = 350$ nm) was probably due to photolysis of trace impurities contained in samples of its precursor, dodecamethylcyclohexasilane. When carefully purified by column chromatography, this material showed only a very weak absorption at the photolysis wavelength used in the experiments (308 nm). Likely impurities which have extremely intense absorption bands in this region are *linear* oligomers or polymers of dimethylsilylene. In view of these uncertainties we wish to withdraw our assignment until a reinvestigation is complete.

Heavy-Atom Kinetic Isotope Effects in the Acid-Catalyzed Rearrangement of *N*-2-Naphthyl-*N'*-phenylhydrazine. Rearrangement Is Shown To Be a Concerted Process [*J. Am. Chem. Soc.* **1985**, *107*, 6674–6678]. HENRY J. SHINE,* LIDIA KUPCZYK-SUBOTKOWSKA, and WITOLD SUBOTKOWSKI

Two errors were made in the treatment of raw data in the calculations of carbon KIE. These errors were in the treatment of "del" values in the ^{13}C isotope-ratio data and in omitting intramolecular competition in the ^{13}C and ^{14}C calculations. Retreatment of "del" data and calculating KIE on the (only available) basis that intra- and intermolecular KIE in the rearrangement of the title compound (**1**) are of equal magnitude gives corrected results for the sought-after intermolecular KIE. The averaged results given in the abstract should be 1.0176 ± 0.0003 for one ^{13}C atom and 1.0287 ± 0.0011 for one ^{14}C atom. These corrections should also appear in the Discussion section of p 6677. The ^{14}C KIE calculated from eq 7 should read (line 32, p 6677) 1.0150 as compared with the experimental result (line 32) 1.0176. The last five entries in column 6, Table II should read in descending order as follows: 1.0252, 1.0259, 1.0281, 0.0174, 1.0162. Similarly, the last five entries in column 7 should read as follows: 1.0284, 1.0276, 1.0302, 1.0178, 1.0173. The transition state for rearrangement of **1** thus, contrary to the earlier indication, is unlike that in the rearrangement of 2,2'-hydrazonaphthalene and more like that in the rearrangement of hydrazobenzene.

Preparation and Characterization of Stable Ruthenium(IV)–Oxo Complexes Which Contain Tertiary Phosphine Ligands [*J. Am. Chem. Soc.* **1986**, *108*, 510–511]. MARY E. MARMION and KENNETH J. TAKEUCHI*

Page 511, left column, final paragraph, fifth sentence: The reported rate constants are incorrect. The sentence should read:

With $[(bpy)_2Ru^{IV}(O)(PPh_3)]^{2+}$ as the catalyst, the oxidation of benzyl alcohol has a second-order rate constant of 1.05 M^{-1} s^{-1}, while the reaction involving $[(bpy)_2Ru^{IV}(O)(PEt_3)]^{2+}$ has a second-order rate constant of 4.2×10^{-3} M^{-1} s^{-1}.

Nucleophilic Attack of Amine and Hydroxide to Platinum Dibenzonitrile Dichloride. Crystal Structure of [$\overline{Pt(NH=\!\!\!=CPhN\text{-}t\text{-}}$ $\overline{BuCH_2CH_2NH\text{-}t\text{-}Bu)Cl(NHCOPh)}$] (2) and *cis*-[Pt(NH$=\!\!\!=$ CPhN-*t*-BuCH₂CH₂NH-*t*-Bu)Cl(NCPh)] (3) [*J. Am. Chem. Soc.* **1986**, *108*, 1180]. LUCIANA MARESCA, GIOVANNI NATILE,* FRANCESCO P. INTINI, FRANCESCO GASPARRINI, ANTONIO TIRIPICCHIO, and MARISA TIRIPICCHIO-CAMELLINI

Page 1181, second full paragraph, 4th row before the end should read: ...absorption at 2280 assignable to $\nu_{C\equiv N}$; other significant bands are at 1565 ($\nu_{C=N}$), ...

Temperature Dependence of the Homolytic Abstraction of Benzylic Hydrogen by Bromine Atom [*J. Am. Chem. Soc.* **1986**, *108*, 2340]. J. G. WRIGHT and R. D. GILLIOM*

Page 2340: The equation shown should be

$$ArCH_3 + \cdot X \rightarrow [ArCH_2H \cdot X \leftrightarrow ArCH_2^+H:X] \rightarrow$$
$$ArCH_2\cdot + HX$$
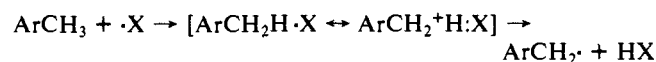
Page 2341: The lengend for Figure 2 should be the following: Hammet plot for the relative rates of substituted toluenes, k', and toluene, k, at -50 °C.

Total Synthesis of (+)-Latrunculin B [*J. Am. Chem. Soc.* **1986**, *108*, 2451–2453]. REGINA ZIBUCK, NIGEL J. LIVERTON, and AMOS B. SMITH III*

Page 2452: structures 16a and 16b are incorrect and should be as illustrated below:
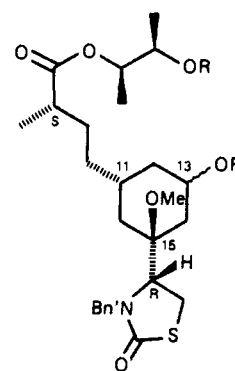
16a R = H
16b R = TBS

Bn' = p-MeOC₆H₄CH₂-

New Synthetic Technology for the Construction of Oxocenes [*J. Am. Chem. Soc.* **1986**, *108*, 2468–2469]. K. C. NICOLAOU,* M. E. DUGGAN, and C.-K. HWANG

Page 2469: footnote 16, should read 200 MHz and not 50.3 MHz; footnote 18 should read ...axial (α) attack... and not ...axial (β) attack....